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stoichiometry? 1. true

2. true 3. false 4. true

5. true 6. 2, 2, 64.10 7.

3, 3, 96.00 8. 2, 2,

88.02 9. 4, 4, 72.08 10.

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methanol and

oxygen gas 11.

carbon dioxide and

water 12. 160.10 g 13.

160.10 g 14. They are

equal. 15. A mole

ratio is a ratio

between the

numbers of moles

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the types of

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indicated by a

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equation. State the

mole ratios from a

balanced chemical

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reactant. the study of quantitative relationships between the amounts of.... in a balanced equation, the ratio between the number of moles.... a reactant that is totally consumed during a chemical reaction....

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371–372 1. Interpret
the following

balanced chemical
equa-tions in terms

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of particles, moles,
and mass. Show that
the law of
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is

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your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2.

3. 4. 5. The study of the quantitative relationships between the

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amounts of reactants
used and the
amounts of products
formed by a chemical
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example, you learned
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reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This

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