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Nomenclature:

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sodium is an example of  
a \_\_\_\_\_ reaction.  
activation energy. in

order for a chemical  
reaction to take place,

the particles invoved  
must collide with a  
sufficient amount of

\_\_\_\_\_

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1) calculate moles of

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each element (if given

mass) 2) divide each  
mole quantity by the

smallest mole value

from step 1. 3) use

whole number values

from step 2 as the

subscripts to the original

equation to write the

empirical formula.

Percent Composition.

-the relative amounts of

each element in a

compound.

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Guide. Molecule. Ionic

Bond. Nonpolar

Covalent Bond.

Polyatomic Ions. unit of

matter that results when

2 or more atoms are

joined by.... A bond

resulting from the

electrostatic attraction

of a cation.... A bond

characterized by an even

distribution of charge;

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(Sections 12-20) The

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length (Section 12

through Section 20) and

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notes and topics

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These guides are closely aligned with chapters 12-17 and 21 and were designed for use in the second semester sequence of general chemistry. They can be used in a flipped-style classroom where students complete them before the lecture.

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length (Section 12 through Section 20) and covers ALL lecture notes and topics discussed in the 2nd semester of your general chemistry lecture course. For the detailed contents

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Thermochemistry  
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(Ch.11): Multiple

Choice 24) What is the amount of heat required to raise the temperature of 200.0 g of aluminum by  $10^{\circ}\text{C}$ ? cal (specific heat of aluminum =  $0.21 \text{ J/g} \cdot ^{\circ}\text{C}$ )  
24) A) 4200 cal  
B) 420 cal  
C) 42 000 cal  
D) 420 000 cal  
E) none of the above

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Chemistry-575 Final

Exam Review Multiple

Choice Identify the

choice that best

completes the statement

or answers the question.

\_\_\_\_\_ 1. How many moles of tungsten atoms are in  $4.8 \times 10^{25}$  atoms of tungsten? a.  $8.0 \times$

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102 moles c.  $1.3 \times 10^1$

moles b.  $8.0 \times 10^1$

moles d.  $1.3 \times 10^2$

moles \_\_\_\_ 2.

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remaining 1/3 will be

composed of matching,

True/False questions,

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Chemistry before 1701  
Chemistry had a bad reputation because of alchemy, which was a pseudoscience that claimed to convert cheap metals into gold.

Alchemy led to the discovery of many elements. 1500 –

George Bauer discovered metallurgy, extracting pure metal out of a mix 1600s –



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Robert Boyle studied  
the behavior of gases  
and is responsible for  
the theory of elements ...

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