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The empirical formula for a compound is C_2H_5 and its relative formula mass is 58.0. Deduce its molecular formula. Deduce its molecular formula. (Relative atomic masses: C = 12.0, H = 1.0)

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formula is NaO, mass is 78

g/mole. b. Empirical formula

is CH₂Cl, mass is 99.0

g/mole. c. Empirical formula

is C₃H₄ mass is 121 g/ mole.

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empirical formula. 2. Divide
the molecular mass by the
empirical mass. 3. Multiply
all of the subscripts in the
empirical formula by the
resulting factor. Example
Problem: What is the
molecular formula for a
material that is 92.2% C and
7.80% H by mass? The
material has a molar mass of
52.0 g/mol. 1. Molecular
mass /

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