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Correct answer: I and III. Explanation: Condition I is always true. Condition II is never true, as Gibbs free energy cannot be negative if enthalpy is positive and entropy is negative. Condition III may be true if temperature is very high (this is the scenario when the term dominates the term).

Thermochemistry and Thermodynamics - Physical Chemistry

Problem : Given that the free energy of formation of liquid water is -237 kJ / mol , calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate ΔG for the reaction, which is $-2 (-237 \text{ kJ / mol}) = 474 \text{ kJ / mol}$. Knowing that $\Delta G = -nFE$ and $n = 4$, we calculate the potential is -1.23 V .

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